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Chapter-10 Examples.
A Carnot vapor
refrigeration cycle is
used to maintain a cold
region at 0.0 F where

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the ambient temperature is 75 °F. Refrigerant R-134a enters the condenser as saturated vapor at 100 lbf/in² and leaves as saturated liquid at the same pressure. The evaporator pressure is 20 lbf/in². The mass flow rate of refrigerant is 12 lbm/s.

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Chapter-10

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Chapter-5 Problems. A rigid tank of volume 10 m^3 contains superheated steam at 1 MPa and $400 \text{ }^\circ\text{C}$. Due to heat loss to the outside atmosphere, the tank gradually cools down to the atmospheric temperature of $25 \text{ }^\circ\text{C}$. Determine (a) the heat transfer and (b) the entropy generated in

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the system's universe
during this cooling
process.

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contents:

thermodynamics .

chapter 01:

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properties and state of
pure substances.

chapter 02: work and
heat. chapter 03:

energy and the first

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law of
thermodynamics.
chapter 04: entropy
and the second law of
thermodynamics.
chapter 05:
irreversibility and
availability

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We discuss the
behavior of a single
subject solving
problems in chemical

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Calculator solves the problem and outputs the solution - specific heat : 232.3 J/(kg*C), quite close to table value for specific heat of silver. $X_{n \times n} = a_{n \times n} \sin^n x$ L # \$ % & ' (n=1) *!

Example Of Problem Solving In Thermodynamics

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52:103 Chemical
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and 1.2 Homework 1
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2.3, 2.4, 2.5 Homework
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2.7, 2.8, 2.15, 2.33
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3

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Problems And

Problem : Calculate the potential of a

concentration cell with anode concentration of 1 M and cathode

concentration of 0.01

M at 75 o C. . Knowing

the Nernst Equation

and realizing that the

temperature is not 25 o

C, we write that: $E = E_o - (RT/nF) \ln Q$

E_o for any concentration cell

is zero so, after

plugging in all the

plugging in all the

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numbers we find that:
 $E = 0.035 \text{ V.}$

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Thermodynamics is an essential subject in the study of the behaviour of gases and vapours in real engineering applications. This book is a complimentary follow up for the book "Engineering Thermodynamics" also

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Basic Concepts And
Definitions 1. A turbine
operating under steady
flow conditions

receives steam at the
following state:

Pressure 13.8bar;

Specific volume 0.143

Internal energy 2590

KJ/Kg; Velocity 30m/s.

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Solution Manual for
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Solution. First we must find the amount of heat released by the ethane. To do this, we calculate the number of moles of ethane gas using the ideal gas equation and multiply the molar heat of combustion by the number of moles. $\Delta H_{\text{combustion}} = 1437.17 \text{ kJ/mol}$

$$n = \frac{PV}{RT}$$

$$n = \frac{0.95 \times$$

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$6.7 \times 10^{-298} \text{ mol}^{-1}$

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MAJORS

Thermodynamic

Properties 1. If an

object has a weight of

10 lbf on the moon,

what would the same

object weigh on

Jupiter? Jupiter 22Moon

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$c \text{ ft} \text{ ft} \text{ lbm-ft} \text{ g} = 75 \text{ g}$

$= 5.4 \text{ g} = 32 \text{ sec} \text{ sec} \text{ lbf-}$

$\text{sec}^2 \text{ c moon} \text{ cmoon}$

Jupiter Jupiter $c \text{ mg}$

$Wg10 \times 32 \text{ W} = m = =$

$59.26 \text{ lb} \text{ gg} 5.4 \text{ mg}$

$59.26 \times 75 \text{ W} = 139 \dots$

Thermodynamic Properties

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advanced

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undergraduate
graduate or graduate
student who already
has a strong
background in
undergraduate
engineering
thermodynamics and is
ready to tackle the
underlying
fundamentals of the
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for those entering
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